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CHM192
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1. Impure solid products of chemical reactions can best be purified by

Distillation

--->> Recrystallization

Fractional distillation

Sublimation

2. The ratio of mass measured in grams per the molar mass of the substance is called

--->> mole

concentration

Normality

Percentage of weight

3. The purity of solid after crystallization can be verified by

Solubility

Ionization

--->> Taking the melting point

Increasing the temperature

4. Calculate the volume of acid needed to prepare 500ml of 2.0 molar concentration of acid, if the stock bottle of the acid reads 1.25 specific gravity, 56% by mass and molar mass of acid 56.

2ml

--->> 8ml

4ml

10ml

5. A solution contains 2.65g of anhydrous sodium carbonate in 250cm³ of solution. Calculate the concentration.

10.6g/dm³

0.1mol/dm³

--->> all of the above

none of the above

6. A standard solution has a known _____.

molar mass

--->> concentration

molarity

normality

7. The most polar solvent used for recrystallization is _____

--->> Water

Ethanol

Diethyl ether

Methanol

8. Gelatinous precipitate such as Aluminium hydroxide are best filtered using

--->> ashless filter paper

sintered glass crucible

ash filter paper

oxalate

9. Which one of the following is not a factor to be considered in selecting a good recrystallization solvent?

--->> Hardness

Flammability

Toxicity

Cost

10. A suitable recrystallization solvent should possess the following features except _____

It should be partially volatile in order to be easily removed from the purified crystals

--->> It should react with the compound being purified.

It should have boiling point below the melting point of compound being purified

It should not react with compound being purified